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dihydrogen (H₂) are reacted over an iron
catalyst to produce ammonia (NH₃):
N₂+3H₂=2NH₃, $\Delta H_r = -109 \text{ kJ/mol}$ (at
773 K) where ΔH_r is the enthalpy of the
reaction (normally referred to as the

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heat of reaction). The Basics of Reaction Kinetics for Chemical Reaction...

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4.8. The rate of the chemical reaction doubles for an increase of 10 K in absolute temperature from 298 K. Calculate E_a . Ans. 4.9. The activation

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energy for the reaction, $2 \text{HI}(\text{g}) \rightarrow \text{H}_2 + \text{I}_2(\text{g})$ is $209.5 \text{ kJ mol}^{-1}$ at 581 K . Calculate the fraction of molecules of reactants having energy equal to or greater than activation energy?

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