

## Chapter 11 The Mole Answer Key

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### Chapter 11 The Mole Answer

312 Chapter 11 The Mole Converting Number of Representative Particles to Moles Zinc is used as a corrosion-resistant coating on iron and steel. It is also an essential trace element in your diet. Calculate the number of moles that contain  $4.50 \times 10^{24}$  atoms of zinc (Zn). 1. You are given the number of atoms of zinc and must find the equivalent number of moles.

### Chapter 11: The Mole

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dividing the mass of the element in one mole of the compound by the mass of one mole of the compound and multiplying by 100. if there is more than one of the element in in the compound like  $\text{CH}_4$  has 4 hydrogen, then you have the multiply the mass of hydrogen by 4 and then divide it by the mass of the  $\text{CH}_4$  and then multiply by 100.

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Episode 11 - The Mole Answer Key 1. Why is it important to use the correct amount of materials in a chemical reaction? If too little is used the reaction may not proceed very far. The use of too much chemical may result in waste. 2. What names are given to the materials at the beginning and end of a chemical reaction? Reactants and products. 3.

### Episode 11 - The Mole

Chapter 10 The Mole Assessment Answers Chapter 11: Stoichiometry A mole ratio is a ratio between the numbers of moles of any two of the substances in a balanced chemical equation For example, consider the reaction shown in Figure 112 In this reaction, potassium (K) reacts with bromine ( $\text{Br}_2$ ) to form potassium bromide

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One mole of carbon still has  $6.022 \times 10^{23}$  carbon atoms, but 98.89% of those atoms are carbon-12, 1.11% are carbon-13, and a trace (about 1 atom in  $10^{12}$ ) are carbon-14. Similarly, the molar mass of uranium is 238.03 g/mol, and the molar mass of iodine is 126.90 g/mol.

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15.2CHAPTER 11: STOICHIOMETRY MOLE TO MOLE RATIO When nitrogen and hydrogen gas are heated under the correct conditions, ammonia gas (NH<sub>3</sub>) is formed.

### **CHAPTER 11: STOICHIOMETRY**

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Mole 1.. Idenü\$-' and calculaterthe number of.representa- five particles. in each of the following quantities. a. 2.15 moles of gold I. b. 0.151 mole ofniü.ogen oxide q.oqxjo NO c. 11.5 moles of potassium bromideG 2. Calculate the number of moles of thè substance that contains the following number of represen- tative parficles.

### **Livingston Public Schools / LPS Homepage**

The mole is the SI unit to measure the amount of a substance. It is the number of representative particles, carbon atoms, in exactly 12g of pure carbon-12. A mole of anything contains  $6.02 \times 10^{23}$  representative particles. A representative particle is any kind of particle such as atoms, molecules, formula units, electrons, or ions.

### **The Mole - Neshaminy School District**

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### **Chemistry Chapter 10 Assessment Answers The Mole**

A mole (mol) is a number of things equal to the number of atoms in exactly 12 g of carbon-12. Experimental measurements have determined that this number is very large:  $1 \text{ mol} = 6.02214179 \times 10^{23}$  things Understand that a mole means a number of things, just like a dozen means a certain number of things—twelve, in the case of a dozen.

### **The Mole - Introductory Chemistry - 1st Canadian Edition**

Chapter 11 - "Like Summer Tempests Came His Tears" Now that Toad is back at the river, he wants to return immediately to Toad Hall. However, Rat has bad news: weasels and stoats (creatures from the Wild Wood) moved into the property when they learned about Toad's sentence, and are now squatting there.

### **The Wind in the Willows Chapters 11 and 12 Summary and ...**

The Mole Chapter Answer Key Answers 1.  $1 \text{ mole} = 6.02 \times 10^{23}$  atoms 2. No, your friend is wrong. The units you will end up with are (atoms)  $2 / \text{mole}$ , which is not what you want. 3.

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